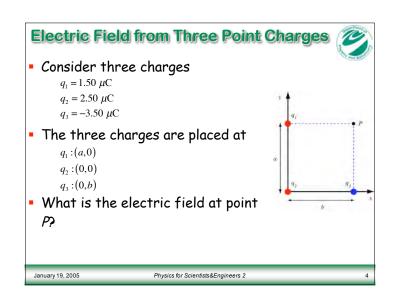
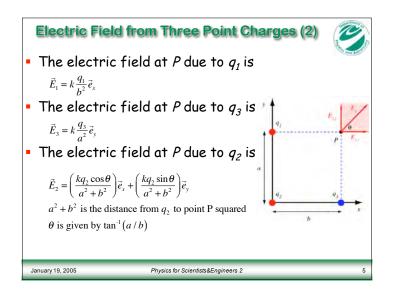
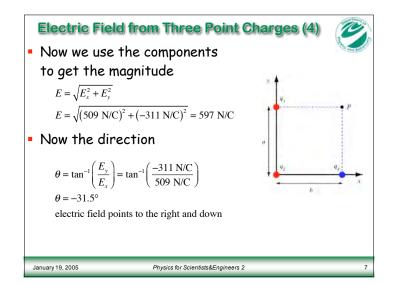
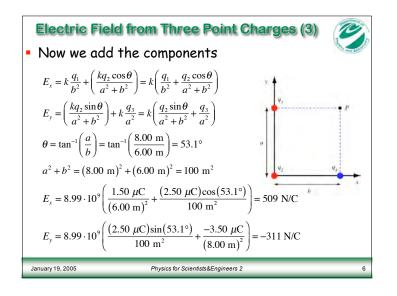


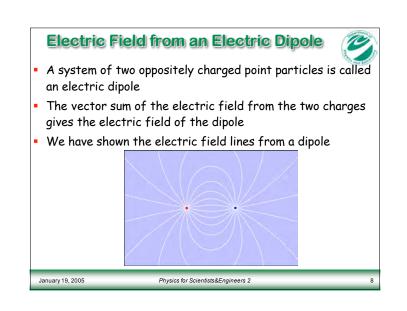
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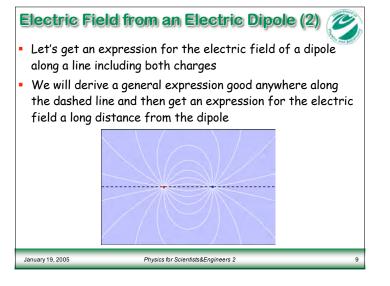












Electric Field from an Electric Dipole (4)

 The principle of superposition tells us that the electric field at any point x is the sum of the electric field from +q and -q

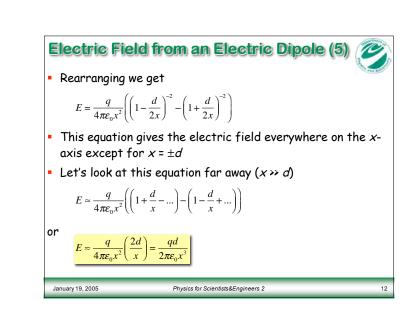
$$E = E_{+} + E_{-} = \frac{1}{4\pi\varepsilon_{0}} \frac{q}{r_{+}^{2}} + \frac{1}{4\pi\varepsilon_{0}} \frac{-q}{r_{-}^{2}}$$

Replacing r₊ and r₋ we get

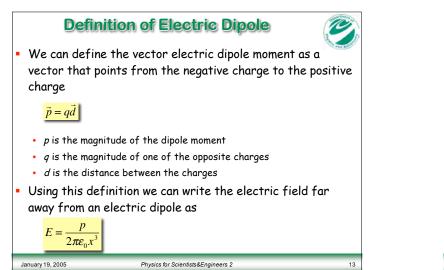
$$E = \frac{1}{4\pi\varepsilon_0} \frac{q}{\left(x - \frac{1}{2}d\right)^2} - \frac{1}{4\pi\varepsilon_0} \frac{q}{\left(x + \frac{1}{2}d\right)^2}$$
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Electric Field from an Electric Dipole (3) Let's put both charges on the x-axis a distance d apart Put -q at x = -d/2 Put +q at x = +d/2 Calculate the electric field at a point P a distance x from the origin x+(d/2) x-(d/2) y +q

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Electric Dipole Moment of Water (2)

- We assume that the center of mass of the two hydrogen atoms is halfway between the two atoms and that the two positive charges are effectively located there
- The distance between the these two positive charges and the two negative charges assumed at the center of the oxygen atom is

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d = (10^{-10} \text{ m})\cos 52.5^\circ = 0.6 \cdot 10^{-10} \text{ m}
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- Our result for the electric dipole moment of water is then $p = 2ed = (3.2 \cdot 10^{-19} \text{ C})(0.6 \cdot 10^{-10} \text{ m}) = 2 \cdot 10^{-29} \text{ C m}$
- This oversimplified result is close to the known value of 6.2 $10^{\text{-30}}$ C \cdot m
 - Assumed charge distribution not what actually happens

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