Quiz on Chapter 12

Quiz 10

1. **C&J page 365 (top), Check Your Understanding #12: "Consider an ob…"**

A) a, b, c, d **B)** b, c, a, d **C)** a, c, b, d **D)** c, b, d, a **E)** b, a, c, d

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2. **Which one of the following statements is the best explanation for the fact that metal pipes that carry water often burst during cold winter months?**

   a) Both the metal and water expand, but water expands to a greater extent.

   b) Freezing water contracts while metal expands at lower temperatures.

   c) The metal contracts to a greater extent than the water.

   d) The interior of the pipe contracts less than the outside of the pipe.

   e) Freezing water expands while metal contracts at lower temperatures.

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3. **25 kJ will melt 4.0 kg of material A. 50 kJ will melt 6.0 kg of material B. 30 kJ will melt 3.0 kg of material C.**
   Rank the heat of fusion of these materials (largest first).

   A) a,b,c **B)** c,b,a **C)** b,c,a **D)** a,c,b **E)** b,a,c

Quiz 10

4. Heat is added to a substance, but its temperature does not increase. Which of the following statements is the best explanation for this?

a) The substance has unusual thermal properties.

b) The substance must be cooler than its environment.

c) The substance must be a gas.

d) The substance must be an imperfect solid.

e) The substance undergoes a change of phase.

Quiz 10

4. Heat is added to a substance, but its temperature does not increase. Which of the following statements is the best explanation for this?

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c) The substance must be a gas.

d) The substance must be an imperfect solid.

e) The substance undergoes a change of phase.

5. What is the final temperature when 2.50×10^5 J are added to 0.950 kg of ice at 0.0 °C (latent heat of fusion for ice, $L_{ice} = 3.35 \times 10^5$ J/kg)

a) 62.8 °C

b) 36.3 °C

c) 15.7 °C

d) 4.2 °C

e) 0.0 °C

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- **a)** 62.8 °C
- **b)** 36.3 °C
- **c)** 15.7 °C
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- **e)** 0.0 °C

To melt all the ice:

$$Q = mL_f = (0.950\,\text{kg})(3.35\times10^5\ \text{J/kg})$$

$$= 3.18\times10^5\ \text{J (don't have this amount)}$$

This much of the ice melts:

$$m_{water} = m_{ice}\left(2.50/3.18\right) = 0.786\ \text{kg}\ \ (\text{at } 0\,°\text{C})$$

This much of the ice remains:

$$m'_{ice} = m_{ice} - m_{water} = \left(0.950 - 0.786\right)\text{kg}$$

$$= 0.164\ \text{kg}\ \ (\text{at } 0\,°\text{C})$$

Chapter 15

Thermodynamics

continued

THE FIRST LAW OF THERMODYNAMICS

The internal energy of a system changes due to heat and work:

$$\Delta U = U_f - U_i = Q - W$$

$Q > 0$ system gains heat

$W > 0$ if system does work

The internal energy (*U*) of an Ideal Gas depends only on the temperature:

Ideal Gas (only): $U = \frac{3}{2} nRT$

$$\Delta U = U_f - U_i = \frac{3}{2} nR(T_f - T_i)$$

Otherwise, values for both Q and W are needed to determine ΔU

Clicker Question 15.1

An insulated container is filled with a mixture of water and ice at zero °C. An electric heating element inside the container is used to add 1680 J of heat to the system while a paddle does 450 J of work by stirring. What is the increase in the internal energy of the ice-water system?

a) 450 J

b) 1230 J

c) 1680 J

d) 2130 J

e) zero J

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a) 450 J

b) 1230 J

c) 1680 J

d) 2130 J

e) zero J

Stirring is work done ON the gas (W is negative)

$$\Delta U = Q - W; \quad W = -450\text{J}$$
$$= (1680 + 450)\ \text{J}$$
$$= 2130\ \text{J}$$

15.4 Thermal Processes

Work done by a gas on the surroundings

$(\Delta P = 0)$ ***isobaric:*** constant pressure: $W = Fs = P(As) = P\Delta V$

$(\Delta V = 0)$ ***isochoric:*** constant volume: $W = P\Delta V = 0$

For an Ideal Gas only

$(\Delta T = 0)$ ***isothermal:*** constant temperature: $W = nRT\ln\left(V_f / V_i\right)$

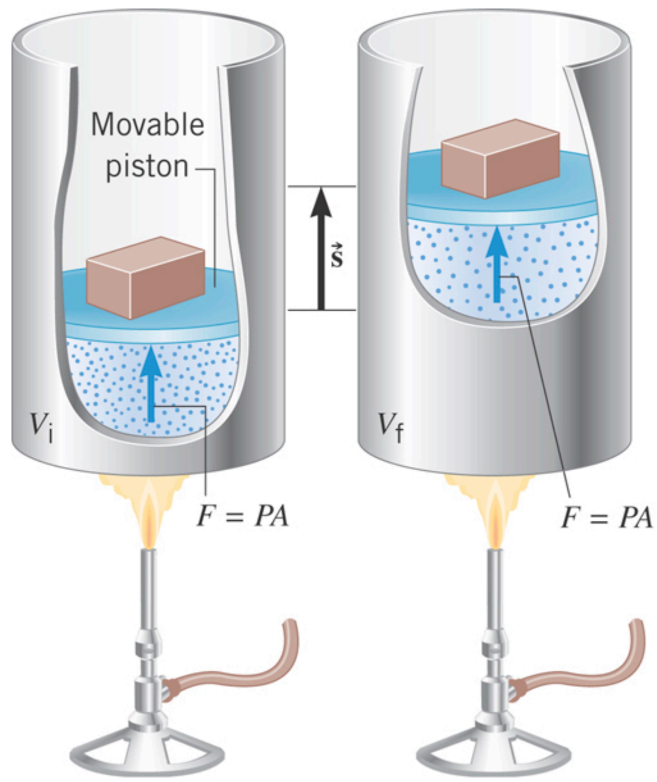
$(Q = 0)$ ***adiabatic:*** no transfer of heat: $W = \frac{3}{2}nR\left(T_f - T_i\right)$

An ***isobaric*** process is one that occurs at constant pressure.

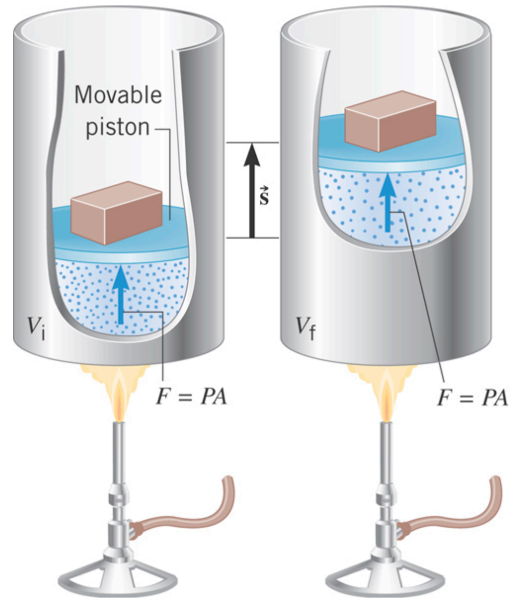
$$W = Fs = P\left(As\right)$$

$$= P\Delta V$$

$$= P\left(V_f - V_i\right)$$



Example 3 Isobaric Expansion of Water (Liquid)



One gram of water is placed in the cylinder and the pressure is maintained at $2.0\text{x}10^5$ Pa. The temperature of the water is raised by 31^{o}C. The water is in the liquid phase and expands by a very small amount, $1.0\text{x}10^{-8}$ m^3.

Find the work done and the change in internal energy.

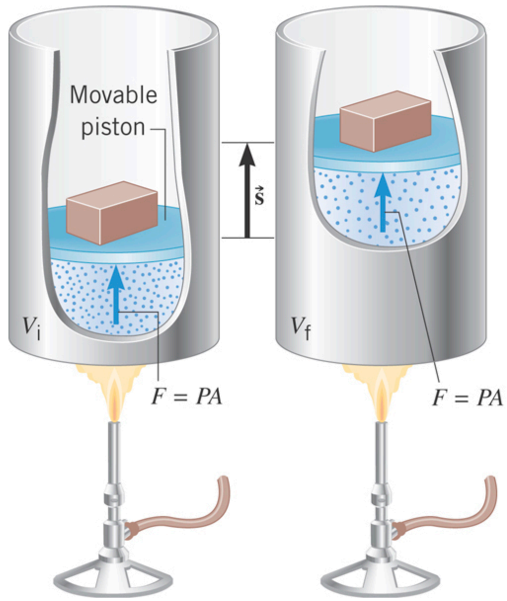
$$W = P\Delta V$$
$$= \left(2.0\times10^{5}\,\text{Pa}\right)\left(1.0\times10^{-8}\,\text{m}^{3}\right) = 0.0020\text{J}$$

$$\text{Liquid water } \Delta V \sim 0$$

$$Q = mc\Delta T$$
$$= \left(0.0010\text{ kg}\right)\left[4186\text{J}/\left(\text{kg}\cdot\text{C}^{\circ}\right)\right]\left(31\text{ C}^{\circ}\right) = 130\text{ J}$$

$$\Delta U = Q - W = 130\text{ J} - 0.0020\text{ J} = 130\text{ J}$$

Example 3 Isobaric Expansion of Water (Vapor)



One gram of water vapor is placed in the cylinder and the pressure is maintained at 2.0×10^5 Pa. The temperature of the vapor is raised by 31°C, and the gas expands by 7.1×10^{-5} m³. Heat capacity of the gas is 2020 J/(kg-C°).

Find the work done and the change in internal energy.

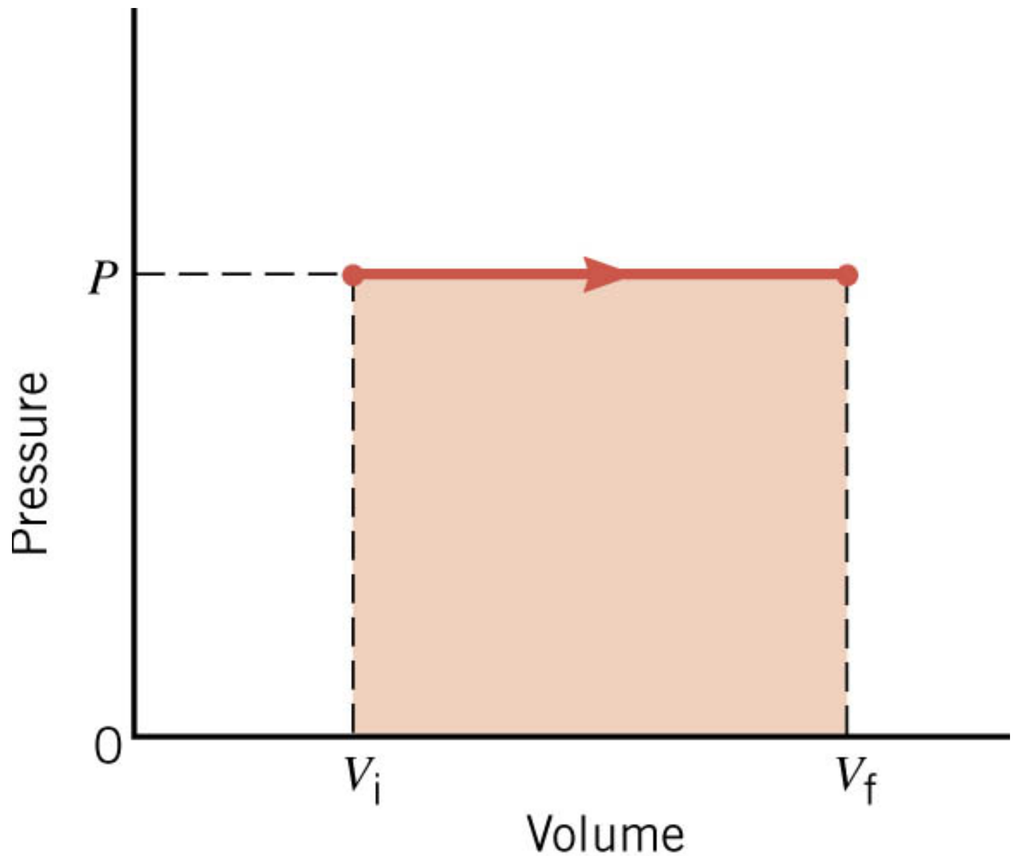
$$W = P\Delta V = (2.0\times10^{5}\ \text{Pa})(7.1\times10^{-5}\text{m}^{3})$$
$$= 14.2\ \text{J}$$

$$Q = mc\Delta T$$
$$= \left(0.0010\ \text{kg}\right)\left[2020\,\text{J}/\left(\text{kg}\cdot\text{C}^{\circ}\right)\right]\left(31\ \text{C}^{\circ}\right) = 63\ \text{J}$$

$$\Delta U = Q - W = 63\ \text{J} - 14\ \text{J} = 49\ \text{J}$$

$$W = P\Delta V = P\left(V_f - V_i\right)$$

The work done at constant pressure the work done is the area under a P-V diagram.



Clicker Question 15.2

An ideal gas at a constant pressure of 1×10^5 Pa is reduced in volume from 1.00 m^3 to 0.25 m^3. What work was done on the gas?

a) zero J

b) 0.25×10^5 J

c) 0.50×10^5 J

d) 0.75×10^5 J

e) 4.00×10^5 J

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c) 0.50×10^5 J

d) 0.75×10^5 J

e) 4.00×10^5 J

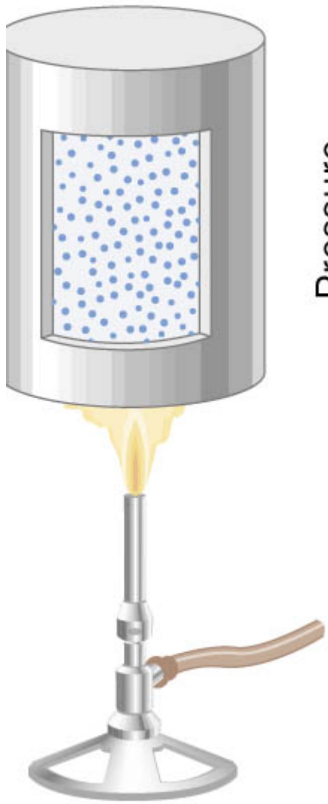
$$W = P\Delta V = P\left(V_f - V_i\right)$$

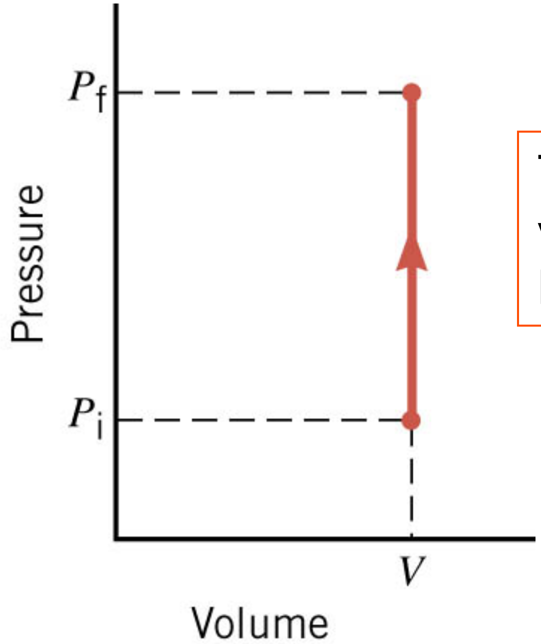
$$= (1 \times 10^5 \text{ Pa})\,(1.00 - 0.25)\text{m}^3$$

$$= 0.75 \times 10^5 \text{ J}$$

15.4 Thermal Processes

isochoric: constant volume


(a)



(b)

The work done at constant volume is the area under a P-V diagram. The area is zero!

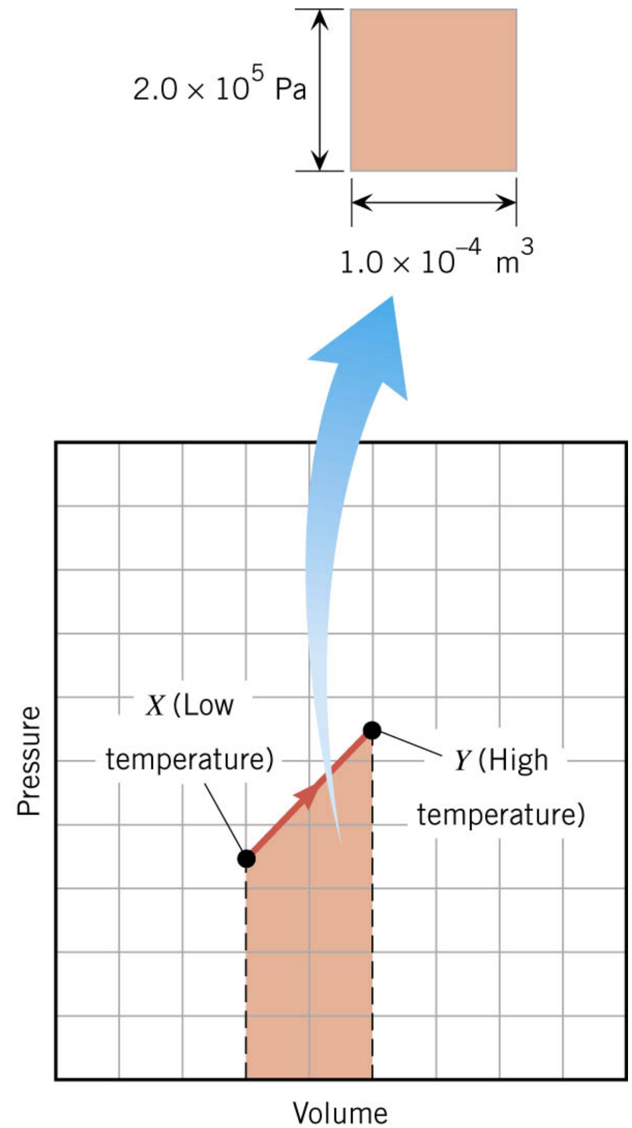
$$W = 0$$

$$\Delta U = Q - W = Q$$

Change in internal energy is equal to the heat added.

Example 4 Work and the Area Under a Pressure-Volume Graph

Determine the work for the process in which the pressure, volume, and temperature of a gas are changed along the straight line in the figure.

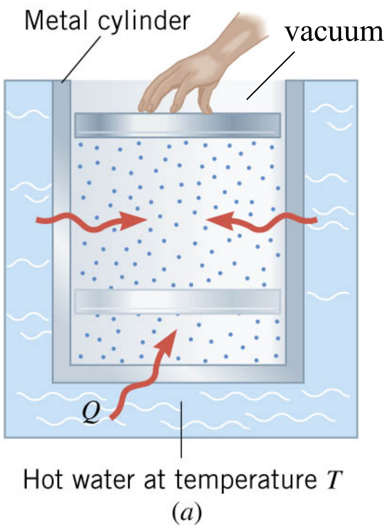


The area under a pressure-volume graph is the work for any kind of process.

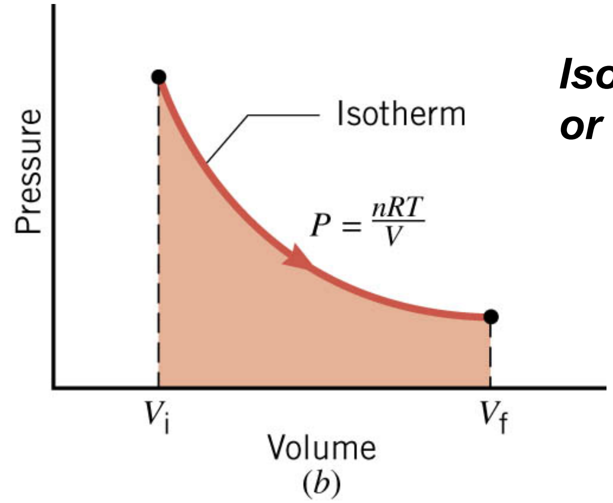
$$W = 9\left(2.0\times10^{5}\,\text{Pa}\right)\left(1.0\times10^{-4}\,\text{m}^{3}\right)$$

$$= +180\ \text{J}$$

15.5 Thermal Processes Using and Ideal Gas



ISOTHERMAL EXPANSION OR COMPRESSION



Isothermal expansion or compression of an ideal gas

$$W = nRT\ln\left(\frac{V_f}{V_i}\right)$$

Example 5 Isothermal Expansion of an Ideal Gas

Two moles of argon (ideal gas) expand isothermally at 298K, from initial volume of 0.025m^3 to a final volume of 0.050m^3. Find (a) the work done by the gas, (b) change in gas internal energy, and (c) the heat supplied.

a) $W = nRT\ln\left(V_f/V_i\right)$

$$= (2.0\ \text{mol})(8.31\text{J}/(\text{mol}\cdot\text{K}))(298\ \text{K})\ln\left(\frac{0.050}{0.025}\right)$$

$$= +3400\ \text{J}$$

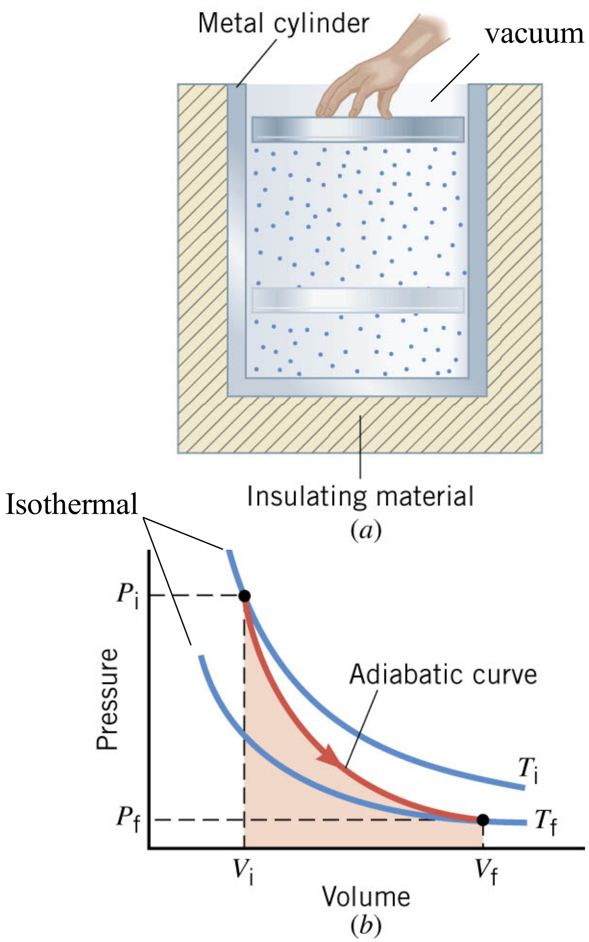
b) $\Delta U = U_f - U_i = \frac{3}{2}nR\Delta T$

$\Delta T = 0$ therefore $\Delta U = 0$

c) $\Delta U = Q - W = 0$

$Q = W = 3400\text{J}$

15.5 Thermal Processes Using and Ideal Gas



ADIABATIC EXPANSION OR COMPRESSION

Adiabatic expansion or compression of a monatomic ideal gas

$$W = \tfrac{3}{2} nR\left(T_i - T_f\right)$$

Adiabatic expansion or compression of a monatomic ideal gas

$$P_i V_i^{\gamma} = P_f V_f^{\gamma}$$

$$\gamma = c_P / c_V$$

Ratio of heat capacity at constant P over heat capacity at constant V.

These are needed to understand basic operation of refrigerators and engines

ADIABATIC EXPANSION OR COMPRESSION

ISOTHERMAL EXPANSION OR COMPRESSION

15.6 Specific Heat Capacities

To relate heat and temperature change in solids and liquids (mass in kg), use:

$$\boxed{Q = mc\Delta T} \qquad \text{specific heat capacity, } c \ \left[\mathrm{J/(kg \cdot {}^\circ C)}\right]$$

For gases, the amount of gas is given in moles, use molar heat capacities:

$$\boxed{Q = nC\Delta T} \qquad \text{molar heat capacity, } C \ \left[\mathrm{J/(mole \cdot {}^\circ C)}\right]$$

$$\boxed{C = (m/n)c = m_u c; \quad m_u = \text{mass/mole (kg)}}$$

ALSO, for gases it is necessary to distinguish between the molar specific heat capacities at constant pressure and at constant volume:

$$C_P, C_V$$

15.6 Specific Heat Capacities

Ideal Gas: $PV = nRT; \quad \Delta U = \frac{3}{2} nR\Delta T$

1st Law of Thermodynamics: $\Delta U = Q - W$

Constant Pressure ($\Delta P = 0$)

$$W_P = P\Delta V = nR\Delta T$$

$$Q_P = \Delta U + W = \frac{3}{2} nR\Delta T + nR\Delta T = \frac{5}{2} nR\Delta T$$

Constant pressure for a monatomic ideal gas

$$Q_P = nC_P \Delta T$$

$$C_P = \frac{5}{2} R$$

Constant Volume ($\Delta V = 0$)

$$W_V = P\Delta V = 0$$

$$Q_V = \Delta U + W = \frac{3}{2} nR\Delta T = \frac{3}{2} nR\Delta T$$

Constant volume for a monatomic ideal gas

$$Q_V = nC_V \Delta T$$

$$C_V = \frac{3}{2} R$$

monatomic ideal gas

$$\gamma = C_P / C_V = \frac{5}{2} R \Big/ \frac{3}{2} R = 5/3$$

any ideal gas

$$C_P - C_V = R$$

15.7 The Second Law of Thermodynamics

The second law is a statement about the natural tendency of heat to flow from hot to cold, whereas the first law deals with energy conservation and focuses on both heat and work.

THE SECOND LAW OF THERMODYNAMICS: THE HEAT FLOW STATEMENT

Heat flows spontaneously from a substance at a higher temperature to a substance at a lower temperature and does not flow spontaneously in the reverse direction.

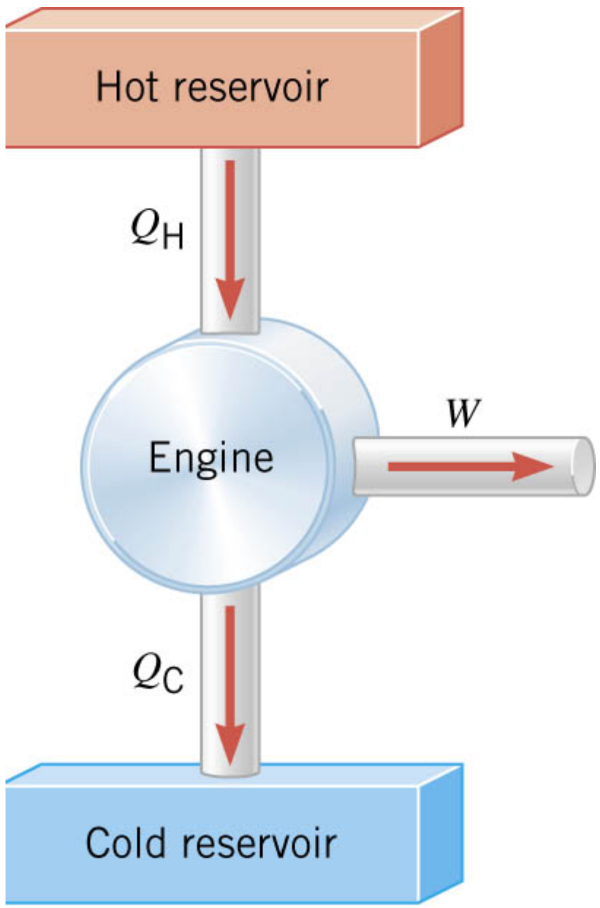
15.8 Heat Engines

A ***heat engine*** is any device that uses heat to perform work. It has three essential features.

1. Heat is supplied to the engine at a relatively high temperature from a place called the *hot reservoir.*

2. Part of the input heat is used to perform work by the *working substance* of the engine.

3. The remainder of the input heat is rejected to a place called the *cold reservoir.*



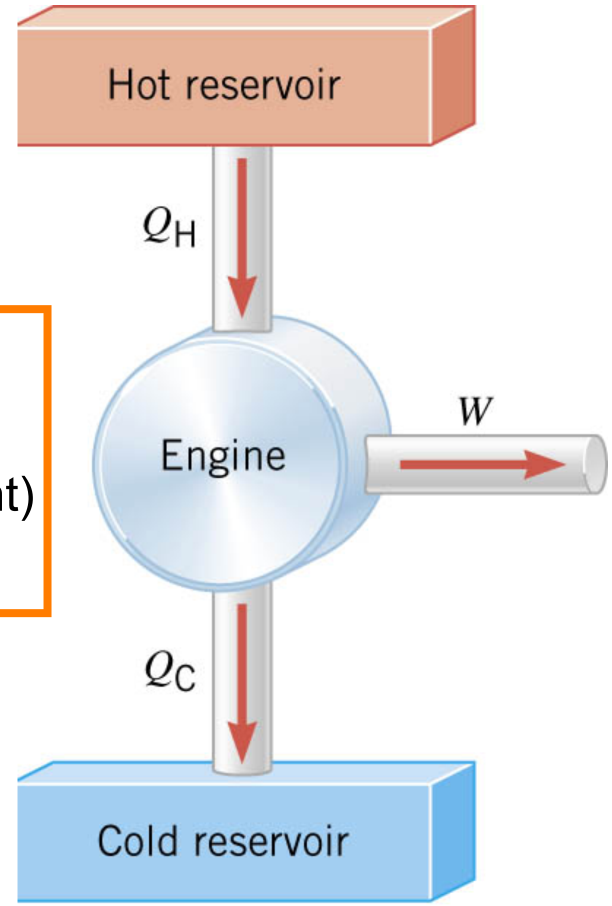
$$\left|Q_H\right| = \text{magnitude of input heat}$$

$$\left|Q_C\right| = \text{magnitude of rejected heat}$$

$$\left|W\right| = \text{magnitude of the work done}$$

Carnot Engine Working with an Ideal Gas

1. ISOTHERMAL EXPANSION ($Q_{in}=Q_H$, T_{Hot} constant)
2. ADIABATIC EXPANSION ($Q=0$, T drops to T_{Cold})
3. ISOTHERMAL COMPRESSION ($Q_{out}=Q_C$, T_{Cold} constant)
4. ADIABATIC COMPRESSION ($Q=0$, T rises to T_{Hot})



$$|Q_H| = \text{magnitude of input heat}$$

$$|Q_C| = \text{magnitude of rejected heat}$$

$$|W| = \text{magnitude of the work done}$$

15.8 Heat Engines

The ***efficiency*** of a heat engine is defined as the ratio of the work done to the input heat:

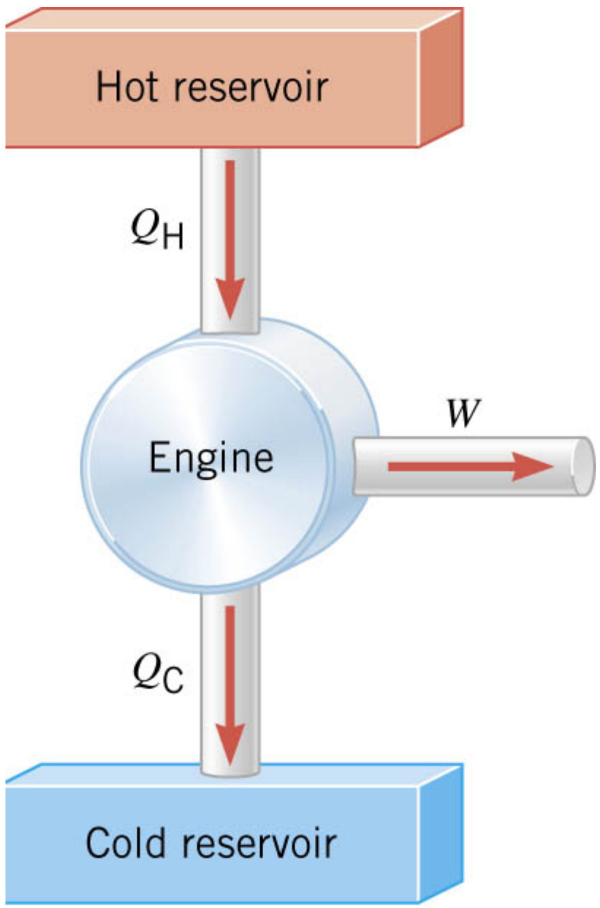
$$e = \frac{|W|}{|Q_H|}$$

If there are no other losses, then

$$|Q_H| = |W| + |Q_C|$$

$$\Downarrow$$

$$e = 1 - \frac{|Q_C|}{|Q_H|}$$



Example 6 An Automobile Engine

An automobile engine has an efficiency of 22.0% and produces 2510 J of work. How much heat is rejected by the engine?

$$e = \frac{|W|}{|Q_H|}$$

$$= \frac{|W|}{|Q_C| + |W|} \quad \Rightarrow \quad e\left(|Q_C| + |W|\right) = |W|$$

$$|Q_C| = \frac{|W| - e|W|}{e} = |W|\left(\frac{1}{e} - 1\right) = 2510\,\mathrm{J}\left(\frac{1}{0.22} - 1\right)$$

$$= 8900\ \mathrm{J}$$

15.9 *Carnot's Principle and the Carnot Engine*

A reversible process is one in which both the system and the environment can be returned to exactly the states they were in before the process occurred.

CARNOT'S PRINCIPLE: AN ALTERNATIVE STATEMENT OF THE SECOND LAW OF THERMODYNAMICS

No irreversible engine operating between two reservoirs at constant temperatures can have a greater efficiency than a reversible engine operating between the same temperatures. Furthermore, all reversible engines operating between the same temperatures have the same efficiency.

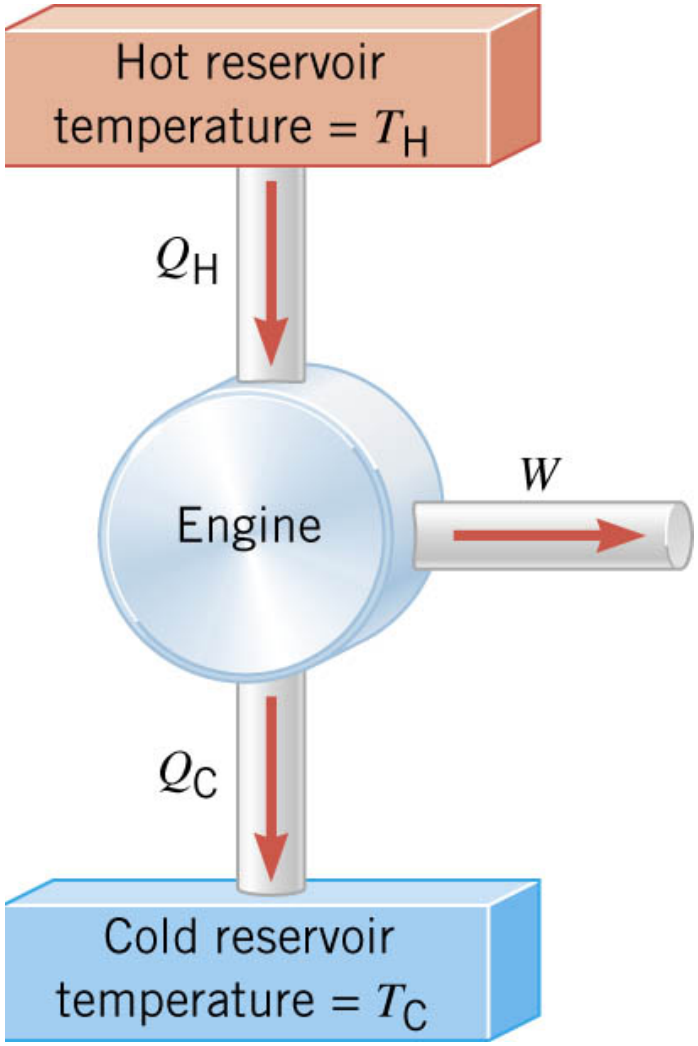
15.9 Carnot's Principle and the Carnot Engine

The ***Carnot engine*** is useful as an idealized model.

All of the heat input originates from a single temperature, and all the rejected heat goes into a cold reservoir at a single temperature.

Since the efficiency can only depend on the reservoir temperatures, the ratio of heats can only depend on those temperatures.

$$e = 1 - \frac{|Q_C|}{|Q_H|} = 1 - \frac{T_C}{T_H}$$



Example 7 A Tropical Ocean as a Heat Engine

Surface temperature is 298.2 K, whereas 700 meters deep, the temperature is 280.2 K. Find the maximum efficiency for an engine operating between these two temperatures.

$$e_{\text{carnot}} = 1 - \frac{T_C}{T_H} = 1 - \frac{280.2\text{ K}}{298.2\text{ K}} = 0.060$$

Maximum of only 6% efficiency.
Real life will be worse.

Conceptual Example 8 Natural Limits on the Efficiency of a Heat Engine

Consider a hypothetical engine that receives 1000 J of heat as input from a hot reservoir and delivers 1000J of work, rejecting no heat to a cold reservoir whose temperature is above 0 K. Decide whether this engine violates the first or second law of thermodynamics.

If $T_H > T_C > 0$

$$e_{\text{carnot}} = 1 - \frac{T_C}{T_H} \text{ must be less than 1}$$

$$e_{hypothetical} = \frac{|W|}{|Q_H|} = \frac{1000\text{ J}}{1000\text{ J}} = 1$$

Violates 2nd law of thermodynamics